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Atmos. Chem. Phys., 8, 4095-4103, 2008

www.atmos-chem-phys.net/8/4095/2008/

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Amines are likely to enhance neutral and ion-induced sulfuric acid-water nucleation in the atmosphere more effectively than ammonia

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Abstract. We have studied the structure and formation thermodynamics of dimer clusters containing H_2SO_4 or HSO_4^- together with ammonia and seven different amines possibly present in the atmosphere, using the high-level ab initio methods RI-MP2 and RI-CC2. As expected from e.g. proton affinity data, the binding of all studied amine- H_2SO_4 complexes is significantly stronger than that of $\text{NH}_3 \cdot \text{H}_2\text{SO}_4$, while most amine- HSO_4^- complexes are only somewhat more strongly bound than $\text{NH}_3 \cdot \text{HSO}_4^-$. Further calculations on larger cluster structures containing dimethylamine or ammonia together with two H_2SO_4 molecules or one H_2SO_4 molecule and one HSO_4^- ion demonstrate that amines, unlike ammonia, significantly assist the growth of not only neutral but also ionic clusters along the H_2SO_4 co-ordinate. A sensitivity analysis indicates that the difference in complexation free energies for amine- and ammonia-containing clusters is large enough to overcome the mass-balance effect caused by the fact that the concentration of amines in the atmosphere is probably 2 or 3 orders of magnitude lower than that of ammonia. This implies that amines might be more important than ammonia in enhancing neutral and especially ion-induced sulfuric acid-water nucleation in the atmosphere.

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Citation: Kurtén, T., Loukonen, V., Vehkamäki, H., and Kulmala, M.: Amines are likely to enhance neutral and ion-induced sulfuric acid-water nucleation in the atmosphere more effectively than ammonia, Atmos. Chem. Phys., 8, 4095-4103, 2008. ▣ [Bibtex](#) ▣ [EndNote](#) ▣ [Reference Manager](#)

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