

# Polysulfide Poisoning of Ag Electrocatalyst during L-Ascorbate Ion Electro-oxidation in Alkaline Solution

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**摘要** L-Ascorbate anion electro-oxidation on a silver electrode in hydroxide solution in the absence and presence of sodium polysulfide of concentrations from  $1 \times 10^{-5}$  to  $4.5 \times 10^{-4}$  mol/L was studied using cyclic voltammetry and electrochemical impedance spectroscopy. Both hydroxide and polysulfide ions inhibited L-ascorbate ion oxidation, with the poisoning effect of polysulfide ion being more pronounced in the potential range of  $-0.3$  to  $-0.2$  V/SCE. The time constants for L-ascorbate ion oxidation in the absence and presence of polysulfide were,  $10^{-3}$  to  $1 \times 10^{-2}$  s and  $1 \times 10^{-4}$  to  $1 \times 10^{-2}$  s, respectively depending on the potential used for the impedance analysis. Based on the cyclic voltammetry findings, a mechanism for L-ascorbate oxidation in the presence of polysulfide ions was proposed. Impedance calculations based on the kinetic analysis can account for the occurrence of a negative impedance in a potential region around  $-0.2$  V/SCE in the Nyquist plots.

**关键词:** [L-ascorbic acid](#) [polysulfide](#) [electrode poisoning](#) [silver](#) [electro-oxidation](#)

**Abstract:** L-Ascorbate anion electro-oxidation on a silver electrode in hydroxide solution in the absence and presence of sodium polysulfide of concentrations from  $1 \times 10^{-5}$  to  $4.5 \times 10^{-4}$  mol/L was studied using cyclic voltammetry and electrochemical impedance spectroscopy. Both hydroxide and polysulfide ions inhibited L-ascorbate ion oxidation, with the poisoning effect of polysulfide ion being more pronounced in the potential range of  $-0.3$  to  $-0.2$  V/SCE. The time constants for L-ascorbate ion oxidation in the absence and presence of polysulfide were,  $10^{-3}$  to  $1 \times 10^{-2}$  s and  $1 \times 10^{-4}$  to  $1 \times 10^{-2}$  s, respectively depending on the potential used for the impedance analysis. Based on the cyclic voltammetry findings, a mechanism for L-ascorbate oxidation in the presence of polysulfide ions was proposed. Impedance calculations based on the kinetic analysis can account for the occurrence of a negative impedance in a potential region around  $-0.2$  V/SCE in the Nyquist plots.

**Keywords:** [L-ascorbic acid](#), [polysulfide](#), [electrode poisoning](#), [silver](#), [electro-oxidation](#)

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- [1] Kumar S A, Cheng H, Chen S. React Funct Polym, 2009, 69: 364 
- [2] Chitravathi S, Kumara Swamy B E, Mamatha G P, Sherigara B S. J Mol Liq, 2011, 160: 193 
- [3] Tian L, Chen L, Liu L, Lu N, Song W B, Xu H D. Sens Ac-tuators B, 2006, 113: 150 
- [4] Fujiwara N, Yasuda K, Ioroi T, Siroma Z, Miyazaki Y, Koba-yashi T. Electrochem Solid-State Lett, 2003, 6(12): A257 
- [5] Fujiwara N, Siroma Z, Ioroi T, Yasuda K. J Power Sources, 2007, 164: 457 
- [6] Majari Kasmaee L, Gopal F. J Power Sources, 2010, 195: 165 
- [7] Fujiwara N, Yamazaki S, Siroma Z, Ioroi T, Yasuda K. J Power Sources, 2007, 167: 32 
- [8] Uhm S, Choi J, Chung S T, Lee J. Electrochim Acta, 2007, 53: 1731 

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- [9] Fujiwara N, Yamazaki S, Siroma Z, Ioroi T, Yasuda K. *Elec-trochem Commun*, 2006, 8: 720
- [10] Power G P, Ritchie I M, Wylie M T. *Electrochim Acta*, 1981, 26: 1633 
- [11] Shimizu K, Aoki K, Osteryoung R A. *J Electroanal Chem*, 1981, 129: 159 
- [12] Aoki K, Shimizu K, Osteryoung R A. *J Electroanal Chem*, 1981, 129: 171 
- [13] Briss V I, Wright G A. *Electrochim Acta*, 1981, 26: 1809 
- [14] Briss V I, Wright G A. *Electrochim Acta*, 1982, 27: 1 
- [15] Horanyi G, Vertes G. *Electrochim Acta*, 1986, 31: 1663 
- [16] Hepel M, Bruckenstein S, Tang G C. *J Electroanal Chem*, 1989, 261: 389 
- [17] Hatchett D W, Gao X, Catron S W, White H S. *J Phys Chem*, 1996, 100: 331
- [18] Vericat C, Vela M E, Gago J, Salvarezza R C. *Electrochim Acta*, 2004, 49: 3643 
- [19] Dryhurst G, Kadish K, Scheller F, Renneberg R. *Biological Electrochemistry*, Vol 1. New York: Academic press, 1982. 256 
- [20] Davey M W, Van Montagu M, Inze D, Sanmartin M, Kanellis A, Smirnof N, Benzie I J J, Strain J J, Favell D, Fletcher J. *J Sci Food Agric*, 2000, 86: 825
- [21] Schrammel A, Koesling D, Schmidt K, Mayer B. *Cardiovasc Res*, 2000, 47: 602 
- [22] Harris D C. *Quantitative Chemical Analysis*. 7th Ed. New York: W.H. Freeman and Company, 2007

- [1] ZHU Lei, MENG Zeda, TRISHA Ghosh, OH Won-Chun. Hydrothermal Synthesis of Porous Ag<sub>2</sub>S Sensitized TiO<sub>2</sub> Catalysts and Their Photocatalytic Activities in the Visible Light Range[J]. *催化学报*, 2012,33(2): 254-260
- [2] K. KOCI1,\* , K. ZATLOUKALOVA1, L. OBALOVA1, S. KREJCIKOVA2, Z. LACNY1, et al. Wavelength Effect on Photocatalytic Reduction of CO<sub>2</sub> by Ag/TiO<sub>2</sub> Catalyst[J]. *催化学报*, 2011,32(5): 812-815
- [3] M. JAFARIAN;\*;F. GOBAL;M. G. MAHJANI;M. HOSSEINI ALIABADI. 碱性溶液中分子氧在多晶银上的多步还原[J]. *催化学报*, 2010,31(5): 541-546